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How many moles of hydrogen atoms  
are in 1 mole of each of the following  
compounds?  $H_2O$  \_\_\_\_\_  $2$  \_\_\_\_\_  $CH_3$

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OH \_\_\_\_\_ 4 \_\_\_\_\_ (NH<sub>4</sub>)<sub>2</sub>CO<sub>3</sub>

\_\_\_\_\_ 8 \_\_\_\_\_ 2. Empirical and Molecular

Formulas Give the empirical formula

for each compound: N<sub>2</sub>H<sub>2</sub>

\_\_\_\_\_ NH<sub>3</sub> \_\_\_\_\_ BH<sub>3</sub> \_\_\_\_\_ BH<sub>3</sub> \_\_\_\_\_ H<sub>2</sub>

O<sub>2</sub> \_\_\_\_\_ OH ...

Stoichiometry Compositions Study

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Stoichiometry. The atomic ratios in each compound are also the relative number of atomic mass units of its elements. The first example is nitrous oxide ( $N_2O$ ), as shown in Table 1. The relative masses were obtained by multiplying the atomic ratios and

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atomic masses. ... CliffsNotes study guides are written by real teachers and professors, so ...

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reaction represented by the equation  
 $2\text{Na} + 2\text{H}_2\text{O} \rightarrow 2\text{NaOH} + \text{H}_2$ , how many

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Section 11.1 What is stoichiometry? 1. true 2. true 3. false 4. true 5. true 6. 2, 2, 64.10 7. 3, 3, 96.00 8. 2, 2, 88.02 9. 4, 4, 72.08 10. methanol and oxygen gas 11. carbon dioxide and water 12. 160.10 g 13. 160.10 g 14. They are equal. 15. A mole ratio is a ratio



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